

SOLUBLE IONIC COMPOUNDS	INSOLUBLE IONIC COMPOUNDS
1. <b>Group 1A ions (Li<sup>+</sup>, Na<sup>+</sup>, K<sup>+</sup>, etc.) and ammonium ion (NH<sub>4</sub><sup>+</sup>)</b> are soluble.	1. <b>(Hydroxides) OH<sup>-</sup></b> and <b>(Sulfides) S<sup>2-</sup></b> , are insoluble <b>except</b> when with Group 1A ions (Li <sup>+</sup> , Na <sup>+</sup> , K <sup>+</sup> , etc.), ammonium ion (NH <sub>4</sub> <sup>+</sup> ) and Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> .
2. <b>(Nitrates) NO<sub>3</sub><sup>-</sup>, (acetates) CH<sub>3</sub>COO<sup>-</sup> or C<sub>2</sub>H<sub>3</sub>O<sub>2</sub><sup>-</sup>,</b> and most <b>perchlorates (ClO<sub>4</sub><sup>-</sup>)</b> are soluble.	2. <b>(Carbonates) CO<sub>3</sub><sup>2-</sup></b> and <b>(Phosphates) PO<sub>4</sub><sup>3-</sup></b> are insoluble <b>except</b> when with Group 1A ions (Li <sup>+</sup> , Na <sup>+</sup> , K <sup>+</sup> , etc.), ammonium ion (NH <sub>4</sub> <sup>+</sup> ).
3. <b>Cl<sup>-</sup>, Br<sup>-</sup>, and I<sup>-</sup></b> are soluble, <b>except</b> when paired with Ag <sup>+</sup> , Pb <sup>2+</sup> , Cu <sup>+</sup> and Hg <sub>2</sub> <sup>2+</sup> .	
4. <b>(Sulfates) SO<sub>4</sub><sup>2-</sup></b> are soluble, <b>except</b> those of Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> , Ag <sup>+</sup> , and Pb <sup>2+</sup> .	